

The Mole Chapter Answer Key

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The Mole Chapter Answer Key

Answer Key Chapter 10: The Mole 10.1 The Mole Concept Questions 1. Imagine that you have 1 mole of coins, each of which is 1.5 mm thick. If they were placed in a single stack, how tall would the stack be (in km)? If the closest distance between the earth and the moon is 356,400 km, would the coins reach the moon? (Hint: Use the

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Chapter 10 Chemical Quantities91 SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287–296) This section defines the mole and explains how the mole is used to measure matter. It also teaches you how to calculate the mass of a mole of any substance. Measuring Matter (pages 287–289) 1.

SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287–296)

mole of water is the water molecule, the representative particle in a mole of copper is the copper atom, and the representative particle in a mole of sodium chloride is the formula unit. 310

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Chapter 11 The Mole Figure 11-2 The amount of each substance shown is 6.02×10^{23} or one mole of representative particles. The representative particle for

Chapter 11: The Mole

The Mole Chapter Answer Key The Mole Chapter Answer Key Answers 1. 1 mole = 6.02×10^{23} atoms 2. No, your friend is wrong. The units you will end up with are (atoms) $2 / \text{mole}$, which is not what you want. 3. If you know the number of molecules and want to find how many atoms of an element are present. You need to multiply the number molecules by

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Answers 1. a. 2 23 2 23 22 2 3molO x5molAl O 7.5molO 2molAl
O 32g 7.5molO x 240gO molO = = b. 23 23 23 23 23 2molAl O
x5molAl 2.5molAl O 4molAl 75.0g 2.5molAl O x 187.5gAl O molAl
O = = c. 23 23 23 23 86.0gAl SO 1.15molAl SO 75g/mol 4molAl
x1.15molAl SO 2.30molAl 2molAl SO 27.0g 2.30molAlx 62.1g mol
= = = 2. ratio of reactants to product is 1:1:1, so need 0.5 mol of
each reactant. BaCl

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Universe59 SECTION 10.1 Movement and Storage of
Groundwater In your

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If we want to make 2 water molecules, we will need 4 hydrogen
atoms and 2 oxygen atoms. If we want to make 5 molecules of
water, we need 10 hydrogen atoms and 5 oxygen atoms. The

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ratio of atoms we will need to make any number of water molecules is the same: 2 hydrogen atoms to 1 oxygen atom. Figure 4.2. 1 Water Molecules.

4.2: The Mole - Chemistry LibreTexts

The mole is a key unit in chemistry. The molar mass of a substance, ... Answer. One mole of elemental nitrogen reacts with three moles of elemental hydrogen to produce two moles of ammonia. In Chapter 5 "Chemical Reactions and Equations: ...

Chapter 6 - Stoichiometry and the Mole - CHE 105/110 ...

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162 Chemistry: Matter and Change • Chapter 10 Solutions Manual CHAPTER 10 SOLUTIONS MANUAL 10. Explain how a mole is similar to a dozen. The mole is a unit for counting 6.02 $\times 10^{23}$ representative particles. The dozen is used to count 12 items. 11. Apply How does a chemist count the number of particles in a given number of moles of a substance?

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Key Concepts and Summary A convenient amount unit for expressing very large numbers of atoms or molecules is the mole. Experimental measurements have determined the number of entities composing 1 mole of substance to be 6.022×10^{23} , a quantity called Avogadro's number. The mass in grams of 1 mole of substance is its molar mass.

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